**TUẦN 10 - HÓA 9**

**ĐÁP ÁN ÔN HỌC SINH YẾU**

**Câu 1:**

a)

1. Fe + 2HCl → FeCl2 + H2
2. FeCl2 + Mg → MgCl2 + Fe
3. MgCl2 + 2NaOH → Mg(OH)2 + 2NaCl

t0

1. Mg(OH)2 MgO + H2O

b)

t0

1. 4Al + 3O2 2Al2O3
2. Al2O3 + 6HCl → 2AlCl3  +3H2O
3. AlCl3 + 3NaOH → Al(OH)3 + 3NaCl
4. 2Al(OH)3 + 3H2SO4(dd) → Al2(SO4)3  + 3H2O

**Câu 2:**

1. H2SO4(dd) + Zn → ZnSO4 + H2

t0

1. 2Fe + 3Cl2  3FeCl3

t0

1. 3Fe + 2O2 Fe3O4
2. Fe + CuSO4 → FeSO4 + Cu
3. 2Al + 6HCl → 2AlCl3 + 3H2

**Câu 3:**

1. 2Al  **+** 3H2SO4(dd)dư  → Al2(SO4)3  + 3H2

2 3 1 3 mol

0,2 0,3 mol

$$n\_{H\_{2}}= \frac{V}{24,79}=\frac{7,437}{24,79}=0,3 (mol)$$

$n\_{Al}$ $=$ $\frac{0,3×2}{3}=0,2 (mol)$

$$m\_{Al}= n×M=0,2×27=5,4 (g)$$

$$m\_{Cu}=m\_{ℎℎA}−m\_{Al}=9,1−5,4=3,7 (g)$$

$$\%m\_{Al}=\frac{m\_{Al}}{m\_{ℎℎA}}×100\%=\frac{5,4}{9,1}×100\%≈59,34\%$$

$$\%m\_{Cu}=\frac{m\_{Cu}}{m\_{ℎℎA}}×100\%=\frac{3,7}{9,1}×100\%≈40,66\%$$